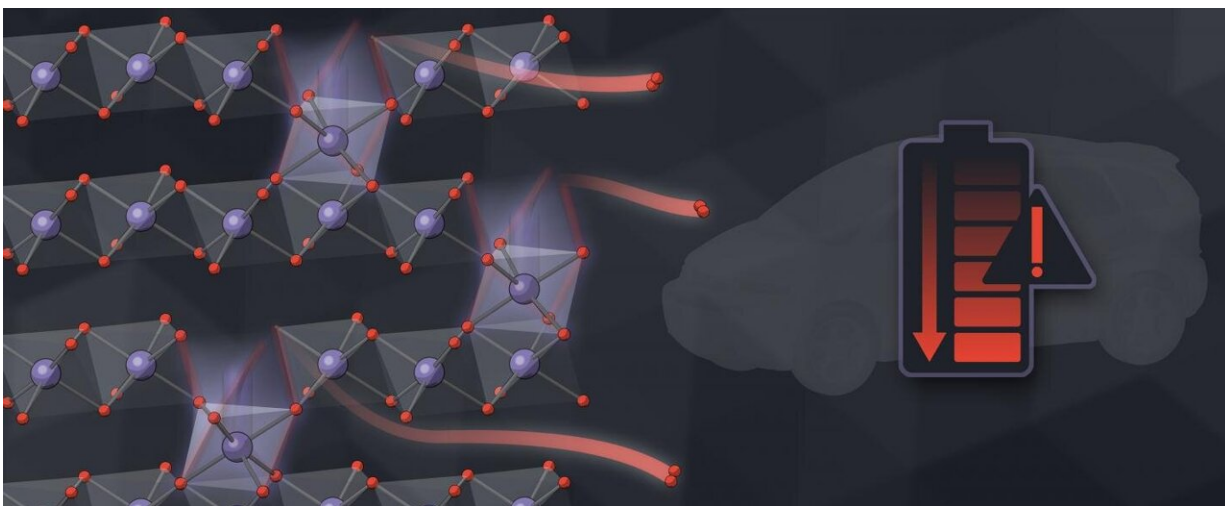


Scientists discover how oxygen loss saps a lithium-ion battery's voltage

June 14 2021, by Glennda Chui



Scientists at SLAC and Stanford have made detailed measurements of how oxygen seeps out of the billions of nanoparticles that make up lithium-ion battery electrodes, degrading the battery's voltage and energy efficiency over time. In this illustration, the pairs of red spheres are escaping oxygen atoms and purple spheres are metal ions. This new understanding could lead to new ways to minimize the problem and improve battery performance. Credit: Greg Stewart/SLAC National Accelerator Laboratory

When lithium ions flow in and out of a battery electrode during charging and discharging, a tiny bit of oxygen seeps out and the battery's voltage—a measure of how much energy it delivers—fades an equally tiny bit. The losses mount over time, and can eventually sap the battery's

energy storage capacity by 10-15%.

Now researchers have measured this super-slow process with unprecedented detail, showing how the holes, or vacancies, left by escaping oxygen atoms change the electrode's structure and chemistry and gradually reduce how much energy it can store.

The results contradict some of the assumptions scientists had made about this process and could suggest new ways of engineering electrodes to prevent it.

The research team from the Department of Energy's SLAC National Accelerator Laboratory and Stanford University described their work in *Nature Energy* today.

"We were able to measure a very tiny degree of oxygen trickling out, ever so slowly, over hundreds of cycles," said Peter Csernica, a Stanford Ph.D. student who worked on the experiments with Associate Professor Will Chueh. "The fact that it's so slow is also what made it hard to detect."

A two-way rocking chair

Lithium-ion batteries work like a rocking chair, moving lithium ions back and forth between two electrodes that temporarily store charge. Ideally, those ions are the only things moving in and out of the billions of nanoparticles that make up each electrode. But researchers have known for some time that oxygen atoms leak out of the particles as lithium moves back and forth. The details have been hard to pin down because the signals from these leaks are too small to measure directly.

"The total amount of oxygen leakage, over 500 cycles of battery charging and discharging, is 6%," Csernica said. "That's not such a small

number, but if you try to measure the amount of oxygen that comes out during each cycle, it's about one one-hundredth of a percent."

In this study, researchers measured the leakage indirectly instead, by looking at how oxygen loss modifies the chemistry and structure of the particles. They tracked the process at several length scales—from the tiniest nanoparticles to clumps of nanoparticles to the full thickness of an electrode.

Because it's so difficult for oxygen atoms to move around in solid materials at the temperatures where batteries operate, the conventional wisdom has been that the oxygen leaks come only from the surfaces of nanoparticles, Chueh said, although this has been up for debate.

To get a closer look at what's happening, the research team cycled batteries for different amounts of time, took them apart, and sliced the electrode nanoparticles for detailed examination at Lawrence Berkeley National Laboratory's Advanced Light Source. There, a specialized X-ray microscope scanned across the samples, making high-res images and probing the chemical makeup of each tiny spot. This information was combined with a computational technique called ptychography to reveal nanoscale details, measured in billionths of a meter.

Meanwhile, at SLAC's Stanford Synchrotron Light Source, the team shot X-rays through entire electrodes to confirm that what they were seeing at the nanoscale level was also true at a much larger scale.

A burst, then a trickle

Comparing the [experimental results](#) with computer models of how oxygen loss might occur, the team concluded that an initial burst of oxygen escapes from the surfaces of particles, followed by a very slow trickle from the interior. Where nanoparticles glommed together to form

larger clumps, those near the center of the clump lost less oxygen than those near the surface.

Another important question, Chueh said, is how the loss of oxygen atoms affects the material they left behind. "That's actually a big mystery," he said. "Imagine the atoms in the nanoparticles are like close-packed spheres. If you keep taking [oxygen atoms](#) out, the whole thing could crash down and densify, because the structure likes to stay closely packed."

Since this aspect of the electrode's structure could not be directly imaged, the scientists again compared other types of experimental observations against computer models of various oxygen loss scenarios. The results indicated that the vacancies do persist—the material does not crash down and densify—and suggest how they contribute to the battery's gradual decline.

"When oxygen leaves, surrounding manganese, nickel and cobalt atoms migrate. All the atoms are dancing out of their ideal positions," Chueh said. "This rearrangement of metal ions, along with chemical changes caused by the missing oxygen, degrades the voltage and efficiency of the battery over time. People have known aspects of this phenomenon for a long time, but the mechanism was unclear."

Now, he said, "we have this scientific, bottom-up understanding" of this important source of battery degradation, which could lead to new ways of mitigating [oxygen](#) loss and its damaging effects.

More information: Peter M. Csernica et al, Persistent and partially mobile oxygen vacancies in Li-rich layered oxides, *Nature Energy* (2021). [DOI: 10.1038/s41560-021-00832-7](https://doi.org/10.1038/s41560-021-00832-7)

Provided by SLAC National Accelerator Laboratory

Citation: Scientists discover how oxygen loss saps a lithium-ion battery's voltage (2021, June 14)
retrieved 9 April 2024 from

<https://techxplore.com/news/2021-06-scientists-oxygen-loss-saps-lithium-ion.html>

This document is subject to copyright. Apart from any fair dealing for the purpose of private study or research, no part may be reproduced without the written permission. The content is provided for information purposes only.