

# Climate-friendly metal production: Why is it difficult?

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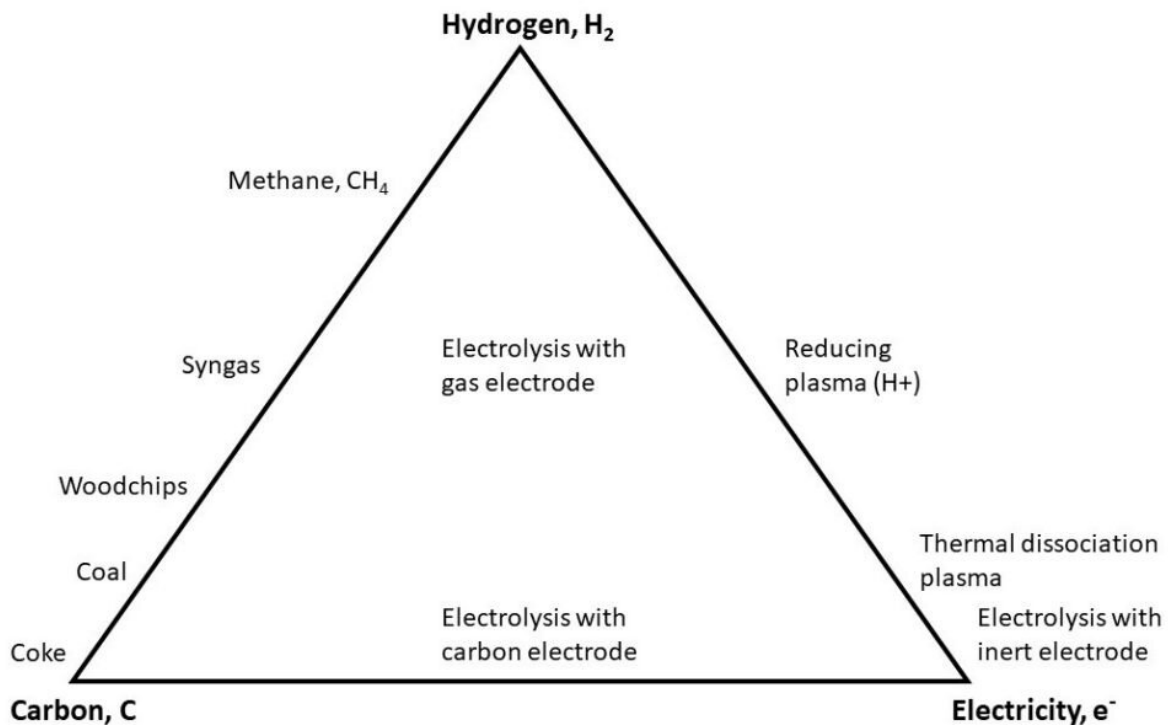


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The figure shows a map of possibilities of which substances that have the ability to remove oxygen from ores, those called reducing agent. Credit: Norwegian University of Science and Technology

Most people know that metals are made from ore, but how do we make

gold from gravel? That's the process we must understand to be able to make the metal industry climate friendly. Here are some alternatives for CO<sub>2</sub>-free metal production.

To achieve "the green shift" the world needs new materials, not least metals. Today, [carbon](#) is particularly popular in the metal industry because it has the ability to remove oxygen from the ore so that we can get the metal out. In that process, CO<sub>2</sub> is formed, the greenhouse gas that we must avoid in the future. At this time, we have started on the task of finding other ways to extract metals, but what alternatives do we really have?

To ensure a future with the development of solar and wind energy, weather-proof roads and cities around the world, we must be able to make metals in a safe, climate-neutral and responsible way. Of course, we must improve the handling, and reusing or recycling of metal-containing waste, but that will not be enough. If we are to achieve the green shift, also globally, the world needs new metals. Today, the metal industry emits [greenhouse gases](#) and in Norway it accounts for 10% of the total CO<sub>2</sub> emissions. It is therefore urgent to find new opportunities for more climate-friendly alternatives to current processes.

## Three keys to future metal production

The figure shows the substances that have the ability to remove oxygen from ores, these are called reducing agents. In the lower left corner of the figure, we see carbon that is used to convert almost all types of ore into metal, today. Historically, the carbon has been charcoal or wood, but in modern times it is mostly fossil carbon that is used. Organic substances such as wood and other forms of biocarbon are not pure forms of carbon, these are found on the line between carbon and [hydrogen](#). Here we also find nontraditional biological carbon sources such as biogas.

Many metal producers see these carbon-based reducing agents as the most interesting reducing agents because they are expected to be able to adapt to current production methods, rather than developing completely new processes. However, the use of all carbon-containing reducing agents will lead to the formation of  $\text{CO}_2$ . In order to prevent this from leading to an increase in the greenhouse effect, one must either use  $\text{CO}_2$ -neutral carbon sources (e.g. biocarbon) or capture and store the  $\text{CO}_2$  from the exhaust gases. If we manage to do both at the same time, we can have  $\text{CO}_2$ -negative processes in the future, which many believe will be necessary to achieve the climate targets. However, the so-called carbon neutral solutions also have a delay, as it takes an average of 90 years for a new tree to grow up and consume the  $\text{CO}_2$  emitted. This is too late if the goals of the Paris Agreement are to be reached.

## **A need for huge amounts of carbon**

The big question regarding carbon is how the access to biocarbon will be like in the future. In addition, we must also think about the impacts on biodiversity and other important environmental aspects. The amounts of carbon needed within the metal sector are, unfortunately, huge. From a long-term perspective, carbon may not even be the best option we can find.

## **Can energy replace reducing agents?**

In the lower right corner of the figure, we have electricity. Many know that aluminum, for example, is made by electrolysis. Then electricity is used to get the oxygen in the ore to let go of the metal. However, today's electrolysis technology lies somewhat towards the carbon corner of the figure, as electrodes made of carbon are needed in the electrolysis. The carbon electrodes are consumed in the process so that one actually combines the effects of electricity and carbon. If we can use other types

of electrodes in the future, and then preferably types that are not used up in the process, the metal could, in theory, be made with only electricity as the reducing agents. Today, large resources around the world are spent on inventing new electrode solutions for different electrolysis processes for different metals. Energy optimists envision a future with large amounts of renewable energy available, and if these forecasts turn out to be correct, such energy-intensive processes will become even more attractive from both an economic and environmental perspective.

If we take a closer look at the electron corner of the figure, we see that there is actually another possibility there: plasma. In fact, if unlimited amounts of energy could be used, it would be possible to make metal with no reduction agent other than clean energy. But then an extreme amount of energy is required, and this option would probably only be possible in the most optimistic energy scenarios.

## **Highly interesting hydrogen**

In the upper corner of the figure, we find hydrogen. Hydrogen is particularly interesting for many reasons, especially since hydrogen is readily available as it is a by-product from several different industrial processes. Unfortunately, it is not possible to use all the hydrogen available today, primarily due to challenges related to transport, storage and safety. Hydrogen can also be made from natural gas, biogas or water (via electrolysis). Unfortunately, however, hydrogen does not have the ability to convert all types of ores into metal, but maybe there are ways to combine hydrogen with other reducing agents to make it more potent?

## **Hydrogen can cooperate**

On the triangle's right edge there is another high-energy alternative: hydrogen plasma. Here, so much energy is added that the hydrogen

atoms have disintegrated. Hydrogen plasma is much more potent as a reducing agent than regular hydrogen gas and can be used on many more ores. This requires more energy than for gas reactions, but considerably less than if plasma is made from the ore itself. Another future possibility revolves around gas electrodes for electrolysis processes. Here it can be envisaged that hydrogen gas is used as a reducing agent in an electrolysis cell.

Another strong candidate for such gas electrodes is methane gas. A gas that is currently most easily obtained from natural gas, but which in the future may come from biological sources, i.e. biogas. Methane can either be a way to add hydrogen to a process, or it can be a reducing agent in itself. The combination of hydrogen and carbon in methane makes this a very interesting option for making metal in the future.

Other gases that can be used to add hydrogen to processes are ammonia, which, like methane, is less explosive than pure hydrogen and therefore easier to transport and store. But using gas as a reducing agent places great demands on restructuring the metal industries, where development and investment of new types of reactors would be necessary.

## **Can metals make metals?**

We should also mention that many metals can be a reducing agent for other metals, although it is not included in the figure here. For example, aluminum can be a reducing agent for many other metals and for silicon. (Silicon is used in solar cells and electronics, for example). The problem with this though, is that you first have to produce aluminum, which currently does not happen without CO<sub>2</sub> emissions. This category of metal production will therefore entirely depend on us putting in place methods both to make and recycle the new reducing agents in a climate-neutral way.

## Why not recycle the carbon?

Speaking of recycling, recycling carbon is a kind of "holy grail." It would be an extremely interesting option, as carbon can be produced without fossil raw materials. At the same time recycling would help reduce the pressure on biological carbon sources, such as forests. If we could capture CO<sub>2</sub> from the exhaust gas and then split it up into oxygen gas, which can be released, and a carbon form that can be put back into furnaces and/or electrolysis cells, it would have been a very attractive solution.

The challenge is that CO<sub>2</sub> is so incredibly stable that enormous amounts of energy will be needed to split it up. The amount of energy can possibly be slightly reduced by using high-tech catalysts, such as Cerium-containing particles or biological organisms such as algae or bacteria.

In other words, there are many different alternatives that all have both advantages and disadvantages. However, there is a good reason to believe that some of these alternatives could be the key to realizing a climate-neutral metal industry of the future.

Provided by Norwegian University of Science and Technology

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