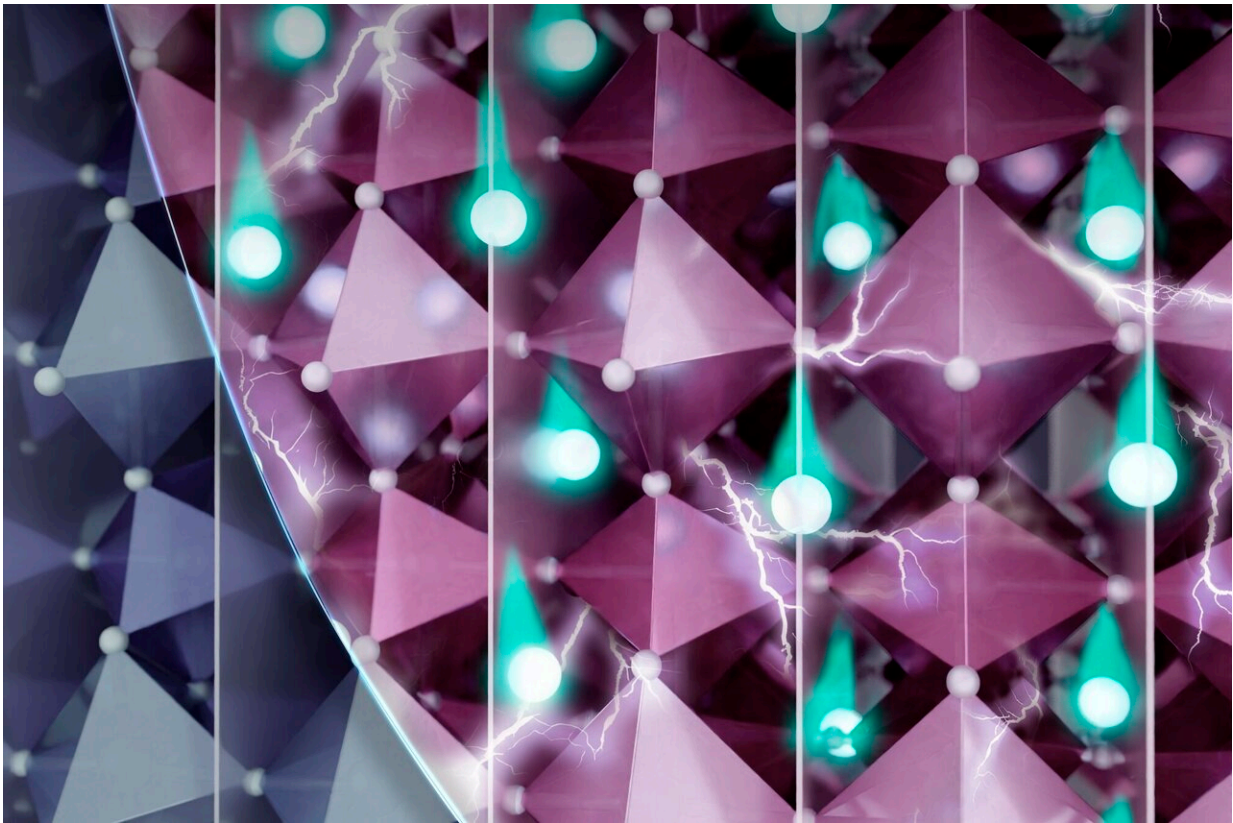


# New research unveils the future of energy-efficient power delivery

July 28 2023, by Nathi Magubane



The rapid movement of lithium ions along the 2D vertical channels in the T-Niobium oxide ( $T\text{-Nb}_2\text{O}_5$ ) thin film results in unique property changes and a chase transition. The blue and purple polyhedra show  $T\text{-Nb}_2\text{O}_5$  lattices, without and with lithium, respectively. The bright green spheres represent lithium ions. Credit: Max Planck Institute of Microstructure Physics / Patricia Bondia

Since the 1940s, scientists have been exploring the use of niobium oxide, specifically a form of niobium oxide known as T-Nb<sub>2</sub>O<sub>5</sub>, to create more efficient batteries. This unique material is known for its ability to allow lithium ions, the tiny, charged particles that make batteries work, to move quickly within it. The faster these lithium ions can move, the faster a battery can be charged.

The challenge, however, has always been to grow this niobium oxide material into thin, flat layers, or "films," that are of high enough quality to be used in practical applications. This problem stems from the complex structure of T-Nb<sub>2</sub>O<sub>5</sub> and the existence of many similar forms, or polymorphs, of niobium oxide.

In a paper published in *Nature Materials*, members of Andrew Rappe's research group at the University of Pennsylvania collaborated with researchers in Max Planck Institute, and University of Cambridge have successfully demonstrated the growth of high-quality, single-crystal layers of T-Nb<sub>2</sub>O<sub>5</sub>, aligned in a way that allows lithium ions to move significantly faster.

"This dramatic shift enables a range of potential applications, from high-speed battery charging to energy-efficient computing and more," says Rappe.

"Our conventional modalities of lithium storage in cathodes usually rely on a process of recrystallization that tends to interfere with the structure, like the ones we see in today's batteries," says co-author Zhen Jiang, a former postdoctoral researcher in the Rappe Group.

Aaron Schankler, a graduate student in the School of Arts & Sciences and adds, "What the team over at Max Planck and the University of Cambridge has done is find a way to move lithium ions in a way that doesn't disrupt the crystal structure of our T-Nb<sub>2</sub>O<sub>5</sub> thin film, and we

helped to rationalize why the ions can move in and out quickly and reversibly."

Rappe likens T-Nb<sub>2</sub>O<sub>5</sub> to a multilevel car-parking structure wherein the lithium ions are cars and the structure of T-Nb<sub>2</sub>O<sub>5</sub> forms open channels, or ramps, that allow the cars to move up and down between the levels.

"By growing the T-Nb<sub>2</sub>O<sub>5</sub> so that these channels run vertically, or 'up and down,' our team made it possible for the lithium ions to move significantly faster, thereby allowing for fast and colossal changes in electrical properties of thin films by inserting the lithium ions between atoms in the negative terminal of our system," says first author Hyeon Han of the Max Planck Institute.

Rappe notes that the researchers from the University of Cambridge worked closely with his team and discovered multiple previously unknown transitions in the material's structure as the concentration of lithium ions changed.

These transitions change the electronic properties of the material, allowing it to switch from being an insulator to a metal, meaning that it goes from blocking electric current to conducting it. This is a dramatic shift; the resistivity of the material decreases by a factor of 100 billion.

The team from Penn developed the computational work in theorizing the conditions needed to give rise to the stability of transitions via density functional theory calculations, a quantum mechanical method used to investigate the electronic structure of many-body systems, especially atoms, molecules, and the condensed phases. Rappe says that with this method the team could calculate and predict material behavior under different conditions.

He says the theoretical calculations helped rationalize the multiple [phase](#)

[transitions](#) they observed, as well as how these phases might be related to the concentration of lithium ions and their arrangement within the crystal structure. This understanding, in turn, allowed the researchers to effectively control and manipulate the electronic properties of the T-Nb<sub>2</sub>O<sub>5</sub> thin films.

"Atomistic simulations calculations have great benefits in advancing the fundamentals of science in academia but also various technologies in industry," says Arvin Kakekhani, a former postdoctoral researcher in the Rappe Group. "This work showcases how these calculations can complement experiments, elucidating the role of [lithium](#) diffusion in electric properties of important solid-state batteries and electronic materials."

"The ability to control the orientation of these films allows us to explore directionally dependent transport in this technologically important class of materials, which is fundamental to understanding of how these materials operate," says Clare P. Grey of the University of Cambridge.

By manipulating these phase transitions, the researchers demonstrated that they could repeatedly and reliably control the electronic properties of these thin films. In addition, by altering the chemical composition of the "gate" electrode, a component that controls the flow of ions in a device, they were able to tune the voltage at which the material becomes metallic, further extending the potential applications.

"This research is a testament to the power of interdisciplinary collaborations and unquenchable scientific curiosity," says Stuart S. P. Parkin of the Max Planck Institute. "Our understanding of T-Nb<sub>2</sub>O<sub>5</sub> and similar complex materials has been dramatically enhanced, setting the course for a more sustainable and efficient future."

**More information:** Hyeon Han et al, Li iontronics in single-crystalline

T-Nb<sub>2</sub>O<sub>5</sub> thin films with vertical ionic transport channels, *Nature Materials* (2023). [DOI: 10.1038/s41563-023-01612-2](https://doi.org/10.1038/s41563-023-01612-2)

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