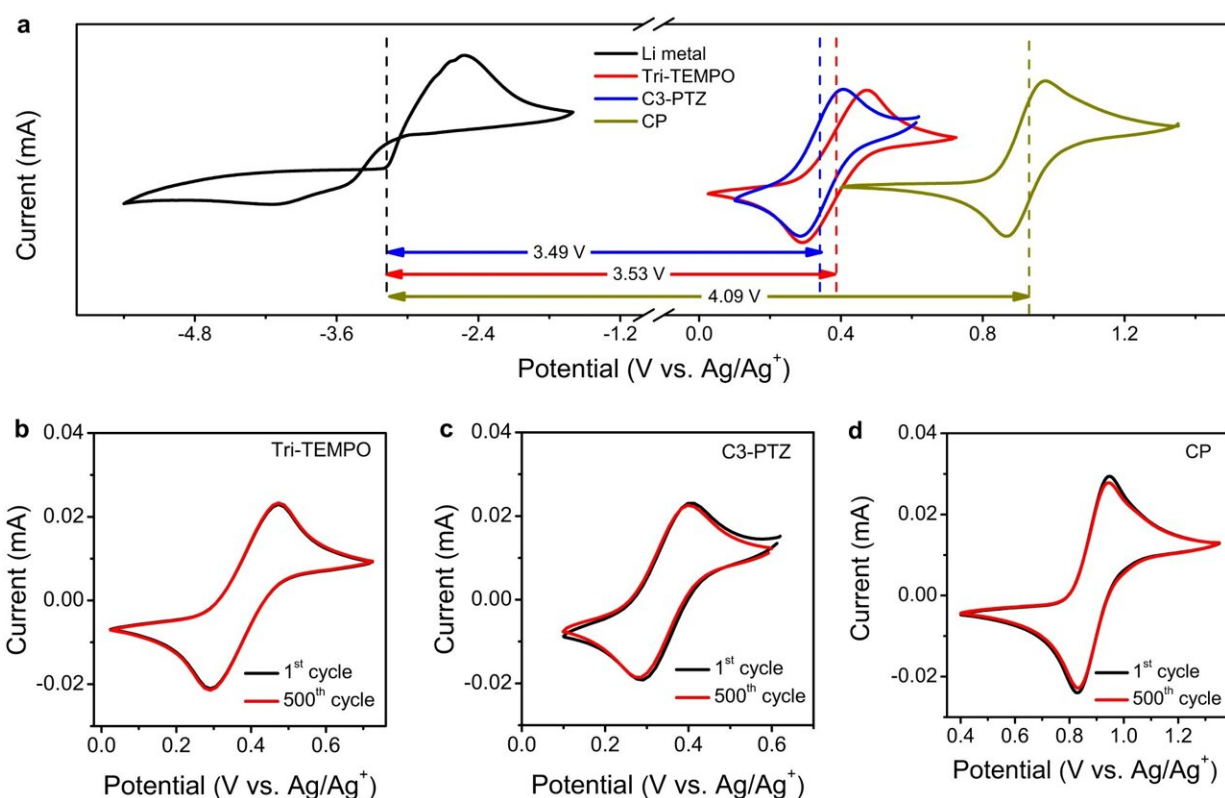


Chemists develop more efficient battery designed for storing energy from wind or solar farms

September 6 2023, by Michael Miller



Electrochemical characterization of redox-active materials. **a** Cyclic voltammograms of Li metal in BMP-TFSI/LiTFSI (0.1 M) (with a Li-metal working electrode, a platinum counter electrode and Ag/Ag⁺ a reference electrode) and redox-active cathode materials including Tri-TEMPO (5 mM), C3-PTZ (5 mM), or CP (5 mM) in FEC/LiClO₄ (0.1 M) solutions (with a glassy carbon (GC) working electrode, a platinum counter electrode and Ag/Ag⁺ a reference electrode). **b–d** Cyclic voltammetry stability of Tri-TEMPO, C3-PTZ,

and CP in FEC/LiClO₄ (0.1 M) solutions. All the CV measurements were conducted under an inert argon atmosphere (at 27 °C) at a scan rate of 50 mV/s. Credit: *Nature Communications* (2023). DOI: 10.1038/s41467-023-40374-y

Jimmy Jiang envisions a future where every house is powered by renewable energy stored in batteries—perhaps even those he and his students are designing today.

In his chemistry lab, Jiang and his [students](#) at the University of Cincinnati have created a [new battery](#) that could have [profound implications](#) for the large-scale energy storage needed by wind and solar farms.

Innovations such as UC's will have profound effects on [green energy](#), Jiang said. Batteries store [renewable energy](#) for when it's needed, not just when it's produced. This is crucial for getting the most out of wind and [solar power](#), he said.

"Energy generation and [energy consumption](#) is always mismatched," he said. "That's why it's important to have a device that can store that energy temporarily and release it when it's needed."

They described their [novel design](#) in [the journal Nature Communications](#).

Traditional car batteries contain a mix of sulfuric acid and water. And while they are inexpensive and made from readily available materials, they have severe drawbacks for industrial or large-scale use. They have a very low energy density, which isn't useful for storing megawatts of power needed to power a city.

And they have a low threshold for electrochemical stability. Jiang said

that means they can blow up.

"Water has a voltage limit. Once the voltage of an aqueous battery exceeds the stability window of 1.5 volts, the water can decompose or be split into hydrogen and oxygen, which is explosive," he said.

But Jiang and his students have developed a battery without water that can generate nearly 4 volts of power. Jiang's novel design does so without a membrane-separator, which are among the priciest parts of these kinds of batteries, he said.

"Membranes are super expensive," said Jiang, an associate professor in UC's College of Arts and Sciences. "We developed a new type of energy storage material that improves performance at a lower cost."

Likewise, membranes are inefficient, he said.

"They can't separate the positive and negative sides completely, so there is always crossover," he said.

The group has submitted provisional patent applications, he said.

"There is still a long way to go," Jiang said.

But he said we are hurtling toward a battery revolution in the next 20 years.

"I am confident about that. There is a lot of intense research going into pushing the boundaries of battery performance," he said.

His students are equally enthusiastic. Doctoral student and study co-author Rabin Siwakoti said the battery offers higher energy density.

"So even a small battery can give you more energy," he said.

"We've managed to eliminate the membrane in a battery, which is a huge component of upfront costs. It's as much as 30% of the cost of the battery," co-author and doctoral student Jack McGrath said.

Co-author Soumalya Sinha, a visiting professor at UC, said countries are racing to develop cheaper, more efficient batteries.

"This design significantly decreases material costs," he said. "We're trying to achieve the same performance at a cheaper cost."

More information: Rajeev K. Gautam et al, Development of high-voltage and high-energy membrane-free nonaqueous lithium-based organic redox flow batteries, *Nature Communications* (2023). [DOI: 10.1038/s41467-023-40374-y](https://doi.org/10.1038/s41467-023-40374-y)

Provided by University of Cincinnati

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