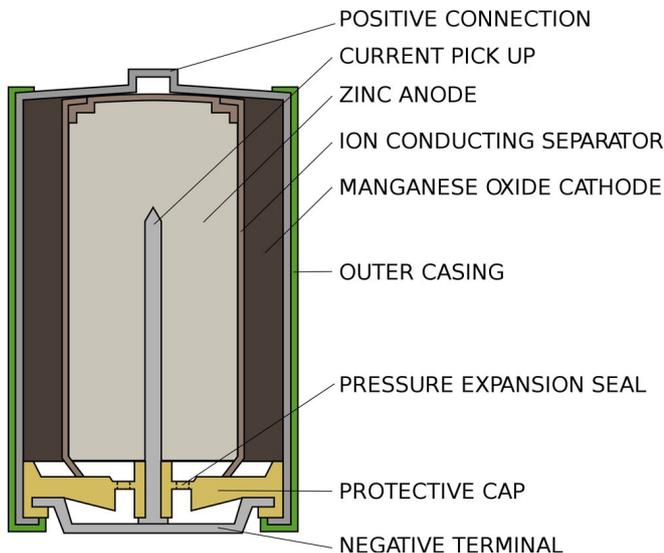


How do lithium-ion batteries work?

15 July 2019, by Robert Masse



What a standard AA alkaline battery looks like on the inside. Credit: [Lead holder/Wikimedia Commons](#)

The smartphone era is only just over a decade old, but the pocket-sized computers at the heart of that societal transformation are only really possible because of another technology: lithium-ion batteries.

First sold commercially [in 1991 by Sony](#) for its camcorders, these types of batteries are good for much more than portable consumer electronics. They're at the center of two other technological revolutions with the power to transform society: the transition from internal combustion engines to [electric vehicles](#), and the shift from an [electric grid](#) powered by fossil fuels to renewable energy generators that store surplus electricity in batteries for future use.

So how do these batteries work? Scientists and engineers have spent entire careers trying to build better batteries and there are still mysteries that we don't fully understand. Improving batteries requires chemists and physicists to look at changes on the [atomic level](#), as well as mechanical and electrical

engineers who can design and assemble the [battery](#) packs that power devices. As a materials scientist at the University of Washington and Pacific Northwest National Lab, [my work](#) has helped explore new materials for [lithium-air batteries](#), magnesium batteries and of course [lithium-ion batteries](#).

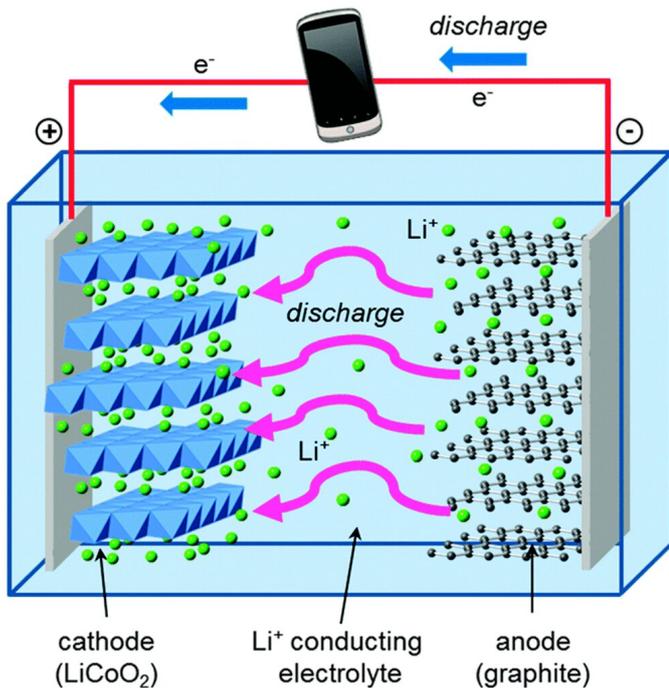
Let's consider a day in the life of two electrons. We'll name one of them [Alex](#) and he has a friend named [George](#).

Battery anatomy

Alex lives inside a standard alkaline AA battery, like in your flashlight or remote control. Inside a AA battery, there is a compartment filled with zinc and another filled with manganese oxide. At one end, the zinc only [weakly hangs onto electrons](#) like Alex. On the other end, the manganese oxide [powerfully pulls](#) electrons toward itself. In between, stopping the electrons from going directly from one side to another, is a piece of paper soaked in a solution of potassium and water, which coexist as positive potassium ions and negative hydroxide ions.

When the battery is put into a device and switched on, the device's internal circuit is completed. Alex gets pulled out of the zinc, through the circuit and into the manganese oxide. Along the way, his movement powers the device, or light bulb or whatever is connected to the battery. When Alex leaves, he can't come back: The zinc that has lost an electron bonds with the hydroxide to form zinc oxide. This compound is extremely stable and cannot easily be converted back into zinc.

On the other side of the battery, the manganese oxide gains an oxygen atom from the water and leaves hydroxide ions behind to balance out the hydroxide being consumed by the zinc. Once all of Alex's neighbors have left the zinc and moved to the manganese oxide, the [battery is exhausted](#) and needs to be recycled.



This process isn't 100% efficient, though—all batteries eventually lose their ability to hold energy. Nevertheless, the family of Li-ion chemistries have been powerful enough to dominate battery technology today.

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As electrons move outside the battery, lithium ions move inside it to keep the electrical equilibrium. Credit: [Islam and Fisher, Chemical Society Reviews, 2014., CC BY](#)

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Lithium-ion advantages

Let's compare this to George, who lives in a lithium-ion battery. Lithium-ion batteries have the same basic building blocks as alkaline AA cells, with a few differences that confer major advantages.

George lives in graphite, which is even weaker than zinc at holding onto electrons. And the other part of his battery is lithium cobalt oxide, which pulls electrons much more powerfully than [manganese oxide](#)—which gives his battery the ability to store much more energy in the same amount of space than an alkaline battery. The solution separating the graphite and lithium cobalt oxide contains positively charged lithium ions, which easily form and break [chemical bonds](#) as the battery is discharged and recharged.

Those [chemical reactions](#) are reversible, unlike the formation of [zinc oxide](#), which is what lets the electrons and the lithium ions flow back and forth over many cycles of charging and discharging.

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